

Name - _____

Start time - __ : __

End time - __ : __

Determine if the statements are true or false.

- Avogadro's number represents the number of atoms in one mole of a substance. (True/False)
- Molar mass is expressed in grams per mole (g/mol). (True/False)
- Stoichiometry involves the study of the mass relationships in chemical reactions. (True/False)
- One mole of any substance always contains the same number of particles. (True/False)

Calculate the molar mass of sulfuric acid (H_2SO_4) given the atomic masses of hydrogen (1 g/mol), sulfur (32 g/mol), and oxygen (16 g/mol).

Discuss the relationship between moles, mass, and Avogadro's number.

Explain why the mole concept is important in chemistry.

a b c d e f g h i j k l m n o p q r s t u v w x y z