

Name - \_\_\_\_\_

Start time - \_\_ : \_\_

End time - \_\_ : \_\_

Determine the formula of the ionic compound formed between the following pairs of ions:

- a.  $\text{Na}^+$  and  $\text{Cl}^-$  \_\_\_\_\_
- b.  $\text{Mg}^{2+}$  and  $\text{O}^{2-}$  \_\_\_\_\_
- c.  $\text{Al}^{3+}$  and  $\text{F}^-$  \_\_\_\_\_

Fill in the blanks with the appropriate terms:

- a. Covalent compounds generally have \_\_\_\_\_ melting points.
- b. They do not conduct electricity in any state because they lack \_\_\_\_\_.
- c. In covalent compounds, atoms share \_\_\_\_\_.
- d. In an ionic compound, electrons are \_\_\_\_\_ from one atom to another.

Explain how an ionic bond forms between a metal and a nonmetal.

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What is the significance of shared electrons in a covalent bond?

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Differentiate between a single bond and a double bond.

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a b c d e f g h i j k l m n o p q r s t u v w x y z